


# C1 REVISION - CHAPTER 4 - CHEMICAL CHANGES - METALS

1) Complete the reactivity series below:

Potassium		_____ reactive
Lithium		
Calcium		
Aluminium		
Iron		
Tin		
Lead		
Silver		
Gold		_____ reactive

2) Explain how the metals in the reactivity series react with water.

3) Explain how the metals in the reactivity series react with acid.

4) What is a displacement reaction?

5) Predict which metals will displace zinc from a solution.

6) Complete the following equation:

Iron + Lead nitrate  $\rightarrow$  \_\_\_\_\_ + \_\_\_\_\_

7) Add hydrogen and carbon to the reactivity series on the left.

Describe how metals are extracted by carbon, using equations.

Describe how metals are extracted by hydrogen, using equations.

Complete the anagram below:-

O  
I  
L  
R  
I  
G

Explain why carbon can reduce zinc oxide but magnesium oxide cannot